Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Period: \_\_\_\_\_ Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Structure of Water and Hydrogen Bonding**

**Getting to the tutorials.**

* Go to [www.sciencemusicvideos.com](http://www.sciencemusicvideos.com); Use the College Bio, AP Bio, or Student Learning Guide Menus to find “The Structure of Water and Hydrogen Bonding”
* *Start with “1. Water Polarity, and Hydrogen Bonds”*

*Note from Mr. W: The videos in these tutorials incorrectly use the term “orbitals” to describe “electron energy levels” (also known as “shells” or “orbits.”). So when you hear “orbital,” think “energy level” or “shell.”*

**Tutorial 1: Water, Polarity, and Hydrogen Bonds.**

**Introduction: Water Makes Life Possible**

1. Most of the mass of any living thing is \_\_\_\_\_\_\_\_\_\_\_\_\_\_.

2. Astronomically, life only exists where water is found in a \_\_\_\_\_\_\_\_\_\_\_\_ form.

**Complete Interactive Reading I.**

3. Complete the table below

|  |  |
| --- | --- |
| Water, Electron Distribution Model | Water, Structural Formula |
|  |  |

4. Why is water the Mickey Mouse molecule?

**Complete Interactive Reading 2: Water and Polarity**

5. Add partial positive and partial negative charges to your diagrams of a water molecule above. (**δ+** and **δ-** for partial positive and negative, respectively).

6. Explain why water has these partial positive and negative charge.

**Complete Interactive Reading 3: Hydrogen Bonding**

7. Draw three or four water molecules, showing hydrogen bonding between them.

8. Explain why water is a polar molecule, and why methane is *not* polar?

9. Explain why water is a liquid at room temperature, while methane is a gas.

10. Take the “Chemistry of Water Quiz.” Check the box when you’re done: ☐

**FINAL REFLECTION:** Watch the video (up at the top of the page. If you’re in class, use ear buds or headphones). Then, use this space to explain why water is sticky, and the consequences of that stickiness.

Proceed to “The Properties of Water.”

**Tutorial 2: The Properties of Water (a virtual lab)**

**Parts 1 and 2: Introduction and Background.**

1. Complete the table below: include partial charges.

|  |  |
| --- | --- |
| Water, structural formula with charges | Alcohol, structural formula with charges |
|  |  |

2. Explain: why is water so much more polar than alcohol?

***Part 3: Six Observations/Experiments Comparing water and alcohol.***

3.1. How is the shape of a drop of water on a penny different from the shape of a drop of alcohol? Explain why it’s different.

3.2. Compare what happens if you try of pile as many drops of water on a penny as you can, and then do the same with alcohol? Why the difference?

3.3. Compare the evaporation speed of water and alcohol. Explain.

3.4. Compare how drops of water and alcohol feel on your skin. Explain.

3.5. Compare what happens when you try to dissolve salt in water with what happens when you try to dissolve salt in alcohol. What’s the difference? Why?

3.6. Compare your attempts to float a paper clip on alcohol and on water? What’s the difference? Why?

4. Use the material in the reading “Five Properties of Water that you should remember” to explain:

a. Why water has a relatively high freezing point.

b. Why ice floats on liquid water (why ice is less dense).

5. Complete the Fill-in-the-Blanks quiz. Check the box when you’re done. ☐

REFLECTION: Watch the video (up at the top of the page). Why is water a stickier molecule than alcohol? How are the properties of each different?

*Click the link to “Acids, Bases and the pH scale*

**Tutorial 3: Acids, Bases and the pH Scale**

1 and 2. Read the Introduction, and Solutions, Acids, and Bases. As you read, write down definitions of the following terms:

a. Solution:

b. Solute:

c. Solvent

Take the Quiz: “:Solutions, Acids, and Bases.” Check the box below when you’re done

☐

Take a minute and write a sentence comparing and contrasting acids and bases. Use sentences like this: While bases are …, acids are….Both are similar in that; they differ because….

3. Read “The Chemistry of Acids and Bases…” When you’re done, check the box below

☐

Copy the formula that shows what happens when HCl is dissolved in water.

Copy the formula that shows what happens when NaOH is dissolved in water.

In a single sentence, summarize how the pH scale works.

4. Take the quiz: Acids, Bases and pH. When you’re done, check the box: ☐

Imagine that someone’s asked you to explain how acids and bases are chemically different: write what you’d say in the space below.

5 and 6. Read the section about buffers, and the conclusion

☐

What are buffers? Why are they important?

REFLECT: What are seven (or six or eight) things that everyone should know about acids, bases, the pH scale, and buffers?

|  |  |
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**Properties of Water, Acids and Bases**



|  |  |
| --- | --- |
| **Across:** | **Down:** |
| 2 - Because of hydrogen bonding, water is a great \_\_\_\_\_\_\_\_ of ionic compounds.6 - A nonpolar compound found in natural gas8 - Nonpolar molecules like methane are a \_\_\_\_\_ at room temperature.10 - Unlike bases, acids will cause \_\_\_\_\_\_\_\_\_\_\_ to dissolve.12 - The acid in your stomach13 - A molecule with regions of positive and negative charge is said to be \_\_\_\_.14 - A hydrogen ion is really just a \_\_\_\_\_\_\_\_\_\_\_\_.15 - Water's ability to form hydrogen bonds makes it a \_\_\_\_\_\_\_\_ molecule.16 - The bond that forms between two water molecules17 - A solution with more hydroxide ions than hydrogen ions.22 - In a water molecule, the \_\_\_\_\_\_\_\_ tend to cluster near the oxygen atom.24 - Because of hydrogen bonding, water has a high heat of \_\_\_\_\_\_\_\_\_\_.26 - A planet with lots of water in liquid form27 - The bonds holding the atoms in a single water molecule together are polar \_\_\_\_\_\_\_ bonds. | 1 - Molecular regions of similar charge will \_\_\_\_\_\_\_ one another.2 - This type of tension exists on the top of a body of water.3 - The first name of a cartoon character reminiscent of water4 - An unusual property of water is that it's less \_\_\_\_\_\_\_ as a solid than as a liquid.5 - Compared to a covalent bond, a hydrogen bond is much \_\_\_\_\_\_\_\_\_\_\_.7 - A water molecule will have a partially \_\_\_\_\_\_\_\_ charge near its oxygen atom.9 - A solution with more hydrogen ions than hydroxide ions11 - By weight, you're mostly made of \_\_\_\_\_\_\_.15 - The thing that gets dissolved in a solution17 - This kind of soda is a weak base18 - The thing that dissolves the solute is a \_\_\_\_\_\_\_\_\_\_\_.19 - A water molecule will have a partially \_\_\_\_\_\_ charge near its hydrogen atoms.20 - The opposite poles of a polar molecule will \_\_\_\_\_\_\_ one another.21 - The mutual attraction between water molecules is called \_\_\_\_\_\_.23 - An acid on a cut will \_\_\_\_\_\_\_\_\_\_\_\_.25 - A water molecule has a single atom of \_\_\_\_\_. |

**Possible Answers:**

Earth, Mickey, acidic, attract, baking, basic, cohesion, covalent, dense, electrons, gas, hydrochloric, hydrogen, metals, methane, negative, oxygen, polar, positive, proton, repel, solute, solvent, solvent, sticky, sting, surface, vaporization, water, weaker